

Sodium in Water Cluster - Part II: Reactivity of a Sodium Atom in Vibrationally Excited Water Clusters



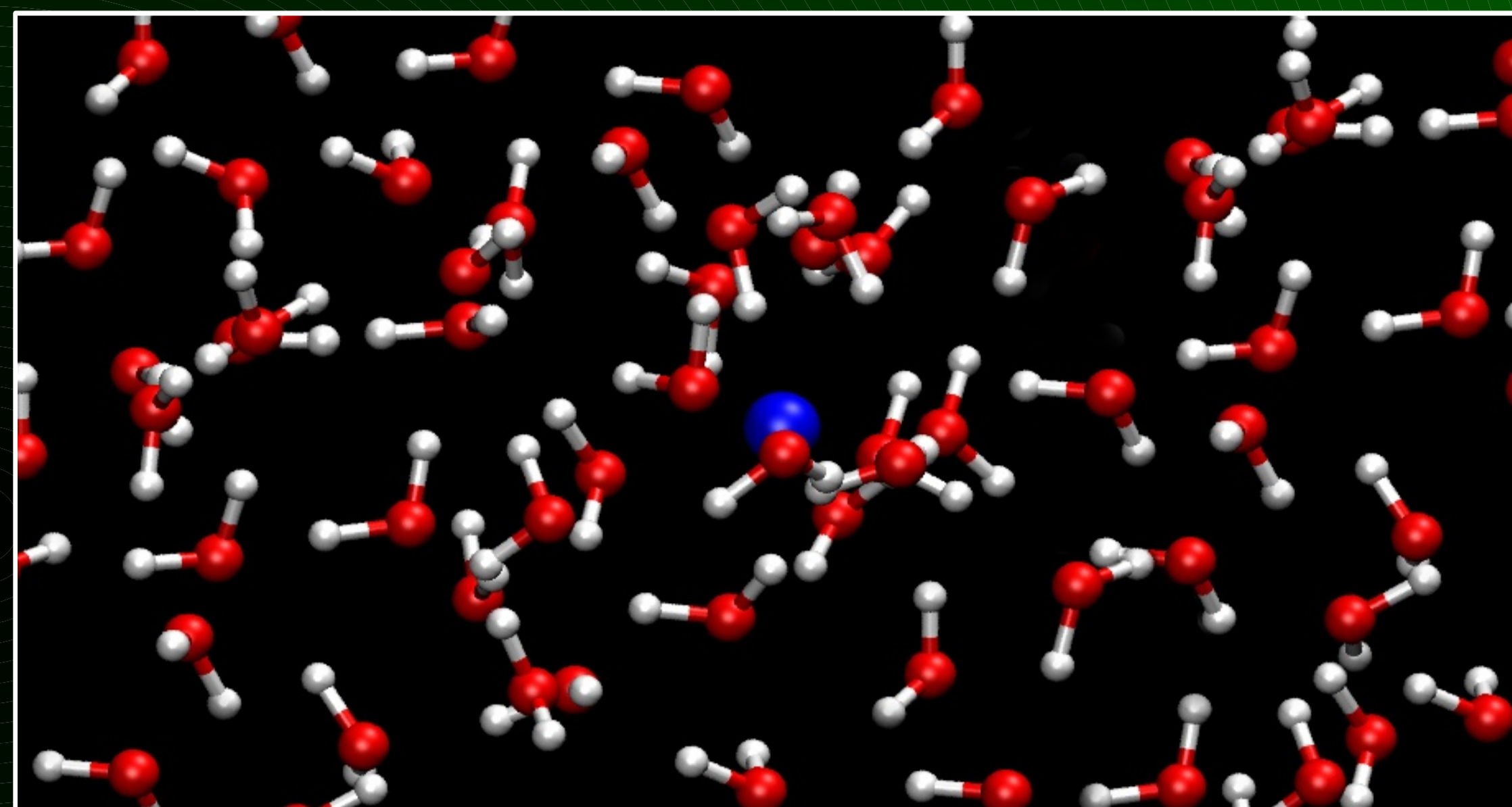
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Big Issue

What is the mechanism of 'sodium-in-water' reaction?

- the reaction is extensively studied and its microscopic mechanism is still an open issue
- recently conducted DFT calculations resulted with proposal of a new one requiring the presence of only one sodium atom^{1,2,3}
- DFT calculations furthermore suggested that the reaction path is connected with an elongation of the O-H bond in a water molecule from the first solvation shell therefore in this work such bonds were vibrationally excited

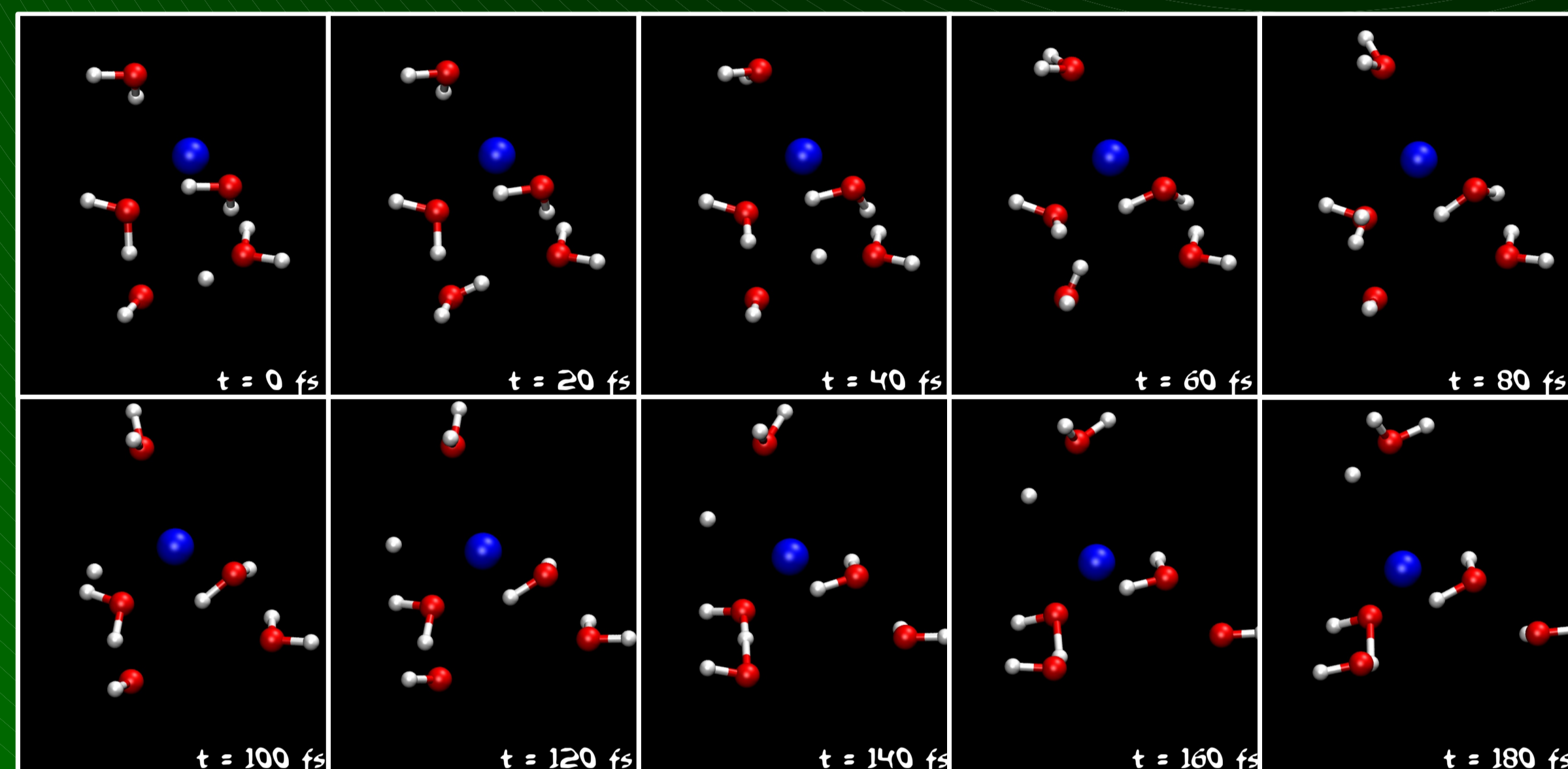


Computational details:

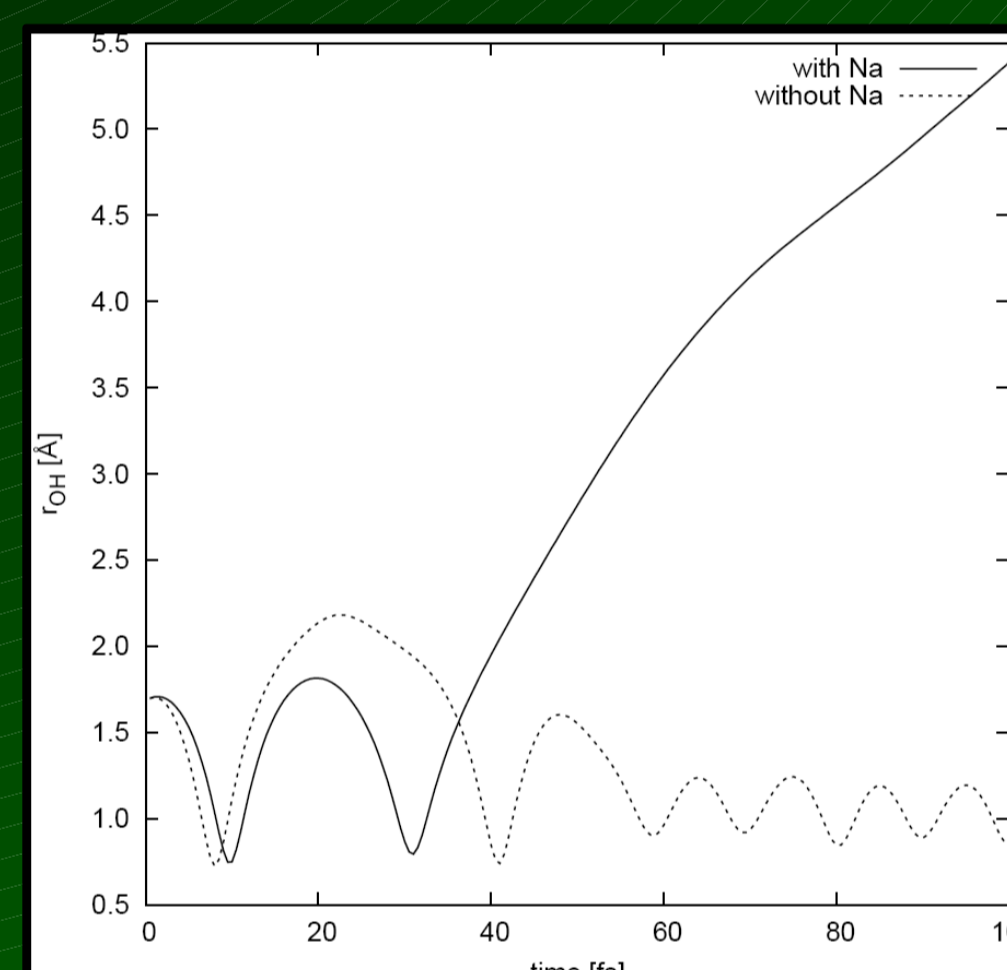
- Born-Oppenheimer Molecular Dynamics from CP2K/Quickstep software package
- Density Functional Theory with mixed Gaussian/PW approach
- BLYP functional with DVZP basis set

Small water clusters (n = 5)

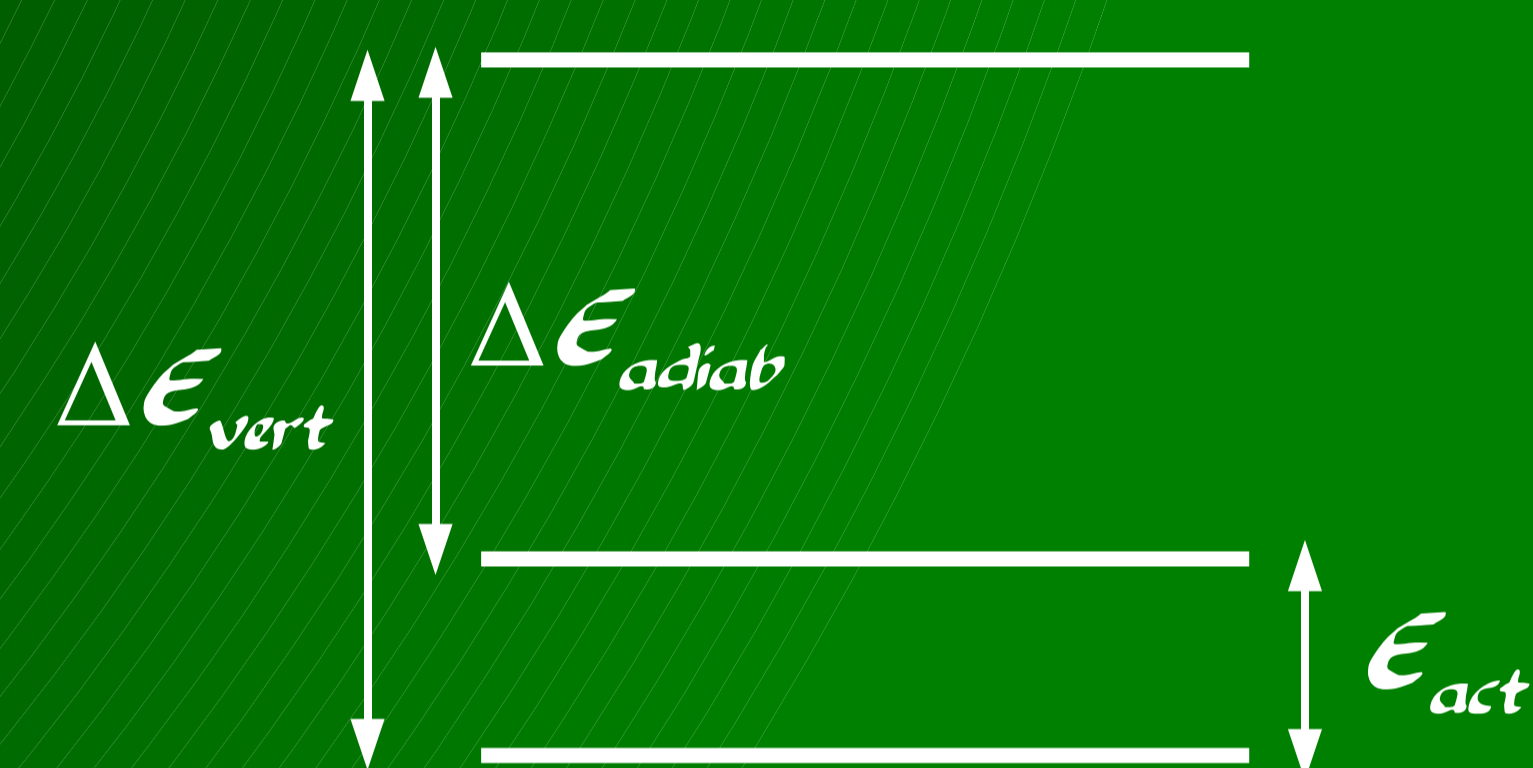
Trajectory



Time evolution of O-H distance (n = 5)



Estimation of activation energy



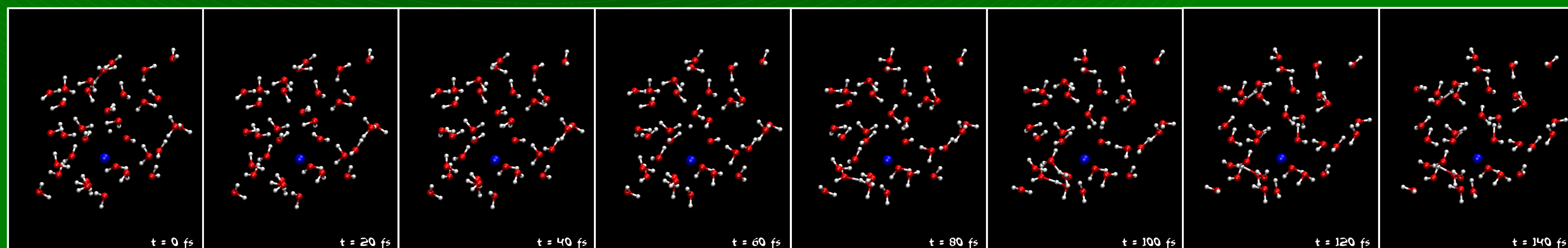
$$\Delta E_{vert} = 83 \text{ kcal/mol}$$

$$\Delta E_{adiab} = 67 \text{ kcal/mol}$$

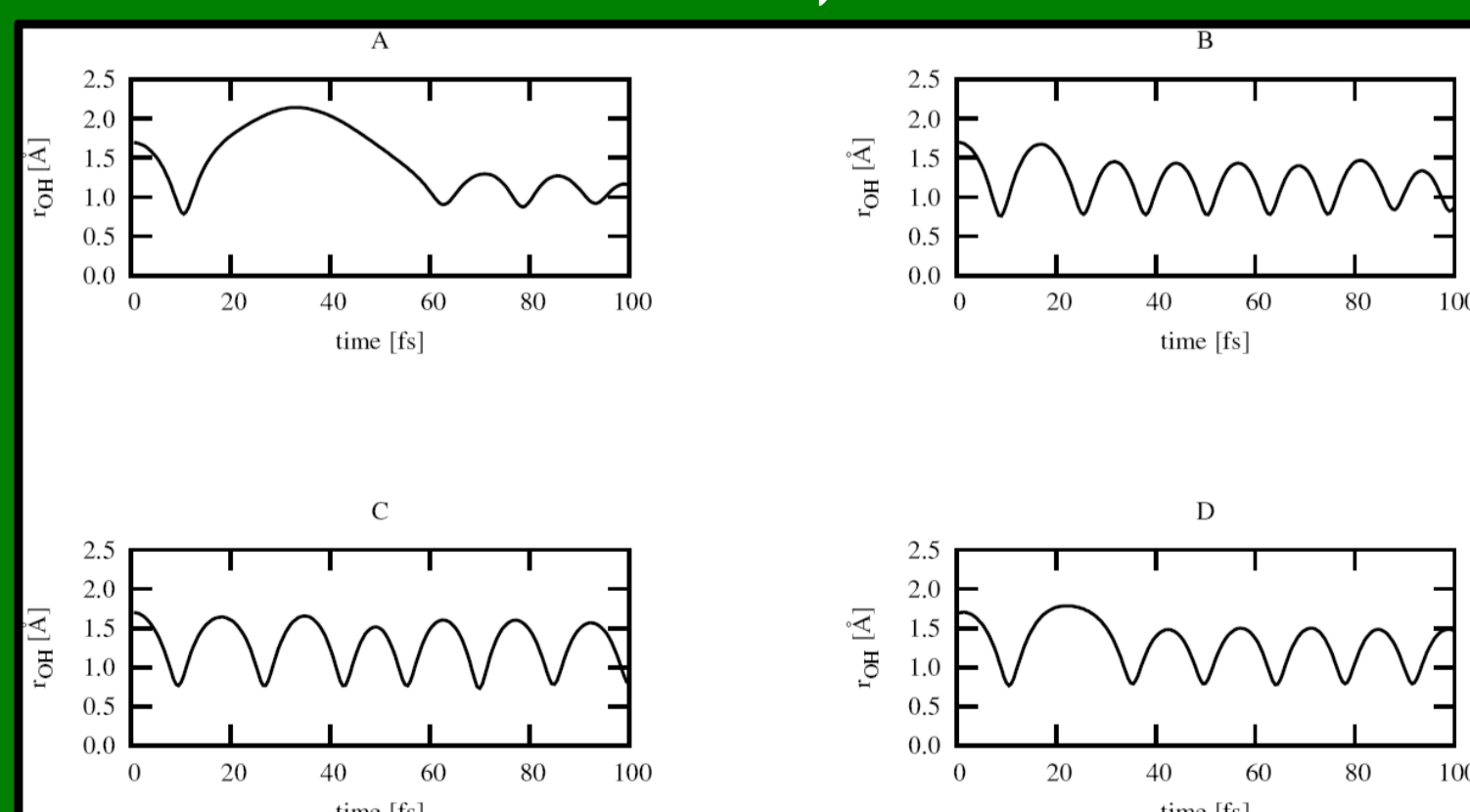
$$E_{act} = \Delta E_{vert} - \Delta E_{adiab} = 16 \text{ kcal/mol}$$

Medium-size water clusters (n = 34)

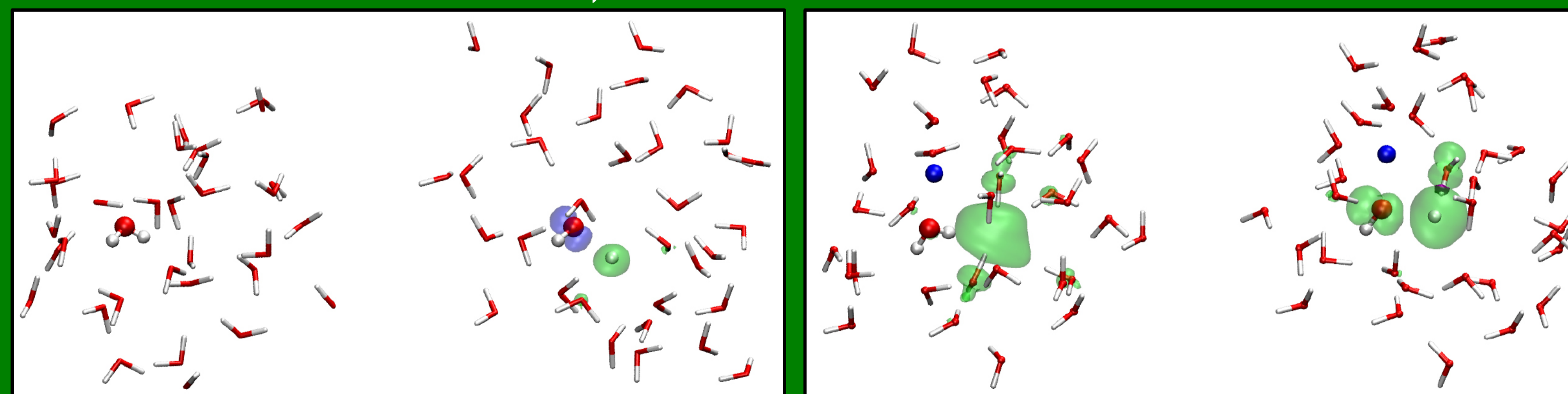
Trajectory



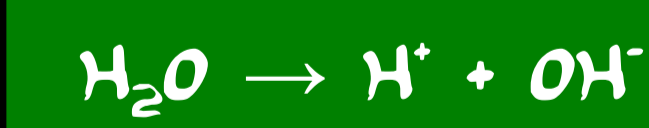
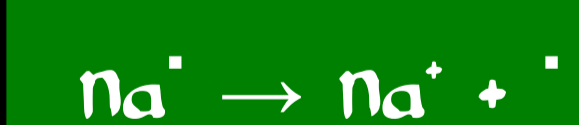
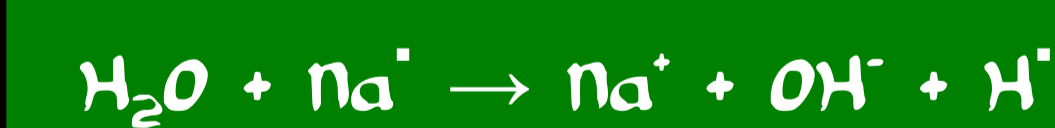
Time evolution of O-H distance



Spin polarization/density for cluster without/with sodium atom



Dissociation reaction



Activation energy is estimated to be less than 16 kcal/mol

Conclusions

- we confirmed the possibility of both water dissociation and NaOH formation in small vibrationally excited water clusters (n = 4, 5)
- vertical vibrational excitation energy of about 80 kcal/mol is needed in order to initiate the reaction
- if water molecules are allowed to relax a low-barrier reaction path is possible (16 kcal/mol)
- we observed water dissociation and a transient Na+OH- formation in a 34 water vibrationally excited cluster

References

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